

5. Match the following compounds to the bond types. Each answer may be used once, more than once, or not at all. C = covalent bond; I = ionic bond; PC = polar covalent.

I a. KBr

PC c. CO₂

PC e. CH₄

C g. Br₂

PC i. SiO₂

I k. NaCl

I b. CuCl

I d. PbO

I f. CaCl₂

I h. Na₂O

I j. ZnI₂

PC l. C₃H₈

6. Complete the following chart by writing the electron configuration notation below each symbol and drawing the Lewis Dot Structure of each element.

H	He
<u>1s¹</u>	<u>1s²</u>
Li	Be
<u>2s¹</u>	<u>2s²</u>
B	C
<u>2s²2p¹</u>	<u>2s²2p²</u>
N	O
<u>2s²2p³</u>	<u>2s²2p⁴</u>
F	Ne
<u>2s²2p⁵</u>	<u>2s²2p⁶</u>
Al	Si
<u>3s²3p¹</u>	<u>3s²3p²</u>
P	S
<u>3s²3p³</u>	<u>3s²3p⁴</u>
Cl	Ar
<u>3s²3p⁵</u>	<u>3s²3p⁶</u>

7. Draw the Lewis Structures for the following molecules.

H ₂ 2	F ₂ 14	H ₂ O 8
NH ₃ 8	C ₂ H ₆ 14	CH ₄ 8